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(Multiply your answer to part a by the kilograms of  
solvent.)  $(1.73\text{m})(0.450\text{kg}) = 0.779\text{ molc}$  Calculate the molar

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mass of the compound. (Take the mass of the solute given in the problem and divide by your answer to part b.)  $135\text{g} \div 0.779\text{mol} = 173\text{ g/mol}$ . Find the molar mass of a covalent compound if 210 g of the substance is dissolved in 810 g of water changes the boiling point of the solution to ...

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Hint: add your answer to question 1 and four times your answer to question 2.  $4 + 4(7) = 4 + 28 = 32$ . Verify that there are 32 electrons pictured in the drawing of  $\text{CCl}_4$ . Count each dot in the picture above: there are 24. There are also 8 more electrons represented by the 4 lines (each line represents two electrons). 5. ChemQuest 23 ...

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51: Cell Voltage - ChemActivity 50: Electrochemical Cell UNIT

11 - HW Practice Keys - ChemQuest 55: Free Energy -

ChemQuest 54: 2nd Law of Thermodynamics

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4. Verify that there are 32 electrons pictured in the drawing of  $\text{CCl}_4$ . Count  
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