

## Answers For Build An Atom Phet Lab

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Responses will vary, but here is what you should see&mldr;. Adding protons changes the identity of the atom adding one moves from right to left on the Periodic Table. You can get the atom to be stable/unstable by adjusting the number of neutrons Adding electrons will change it from a +ion to neutral atom, to ion depending on how many electrons you add Adding electrons you can only add 2 to the first orbital and 8 to the second b) What particle(s) are found in the center of the atom?

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BUILD AN ATOM ATOM SCREEN 1. Explore the Build an Atom simulation. a) List two things you observed in the simulation. Proton, neutron and electron are different colors The amount of these three changes the element b) What particle(s) are found in the center of the atom? Protons and neutrons 1. Play until you discover which particle(s) determine(s) the name of the element you build.

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hydrogen2 and hydrogen3 are both isotopes of hydrogen1 magnesium25 is an isotope of magnesium24 carbon13 and carbon14 are both isotopes of carbon12 (carbon12 1 is an anion of carbon12, carbon12 +1 is a cation of carbon12) anions are negative ions with more electrons than protons cations are positive ions with fewer electrons than protons (the Gallium examples are also ions) 5 BUILD AN ATOM 6 BUILD AN ATOM

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?Build an Atom? - PhET Interactive Simulations

## Access Free Answers For Build An Atom Phet Lab

### ?Build an Atom? - PhET Interactive Simulations

Build an atom out of protons, neutrons, and electrons, and see how the element, charge, and mass change. Then play a game to test your ideas! Sample Learning Goals Use the number of protons, neutrons, and electrons to draw a model of the atom, identify the element, and determine the mass and charge. Predict how addition or subtraction of a ...

### Build an Atom - Atoms | Atomic Structure | Isotope Symbols ...

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“Build An Atom” Simulation – Build an Ion Learning Goals: 1. Draw models that show atomic structure. 2. Use information about the number of protons, neutrons, and electrons to identify an element, its position on the periodic table. 3. Predict how changing the number protons and electrons will change the element and its charge.

### “Build An Atom” Simulation – Build an Ion

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