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Describe how a substance can dissolve in water by dissociation, dispersion, or ionization. Describe how the physical

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properties of a solution can differ from those of its solute and solvent Identify energy changes that occur during the formation of a solution Describe factors affecting the rate at which a solute dissolves in a solvent Vocabulary: solute solvent dissociation dispersion ...

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During the formation of a solution, energy is either absorbed or released. Solutions can become warmer (exothermic) or colder (endothermic) based on this. What factors affect the rate of dissolving? In order for a solvent to form, both the ATTRACTIONS of SOLUTE AND SOLVENT must be broken which takes energy.

8 Solutions, Acids and Bases: 8.1 Formation of Solutions 8 ...

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Section 81 Formation Of Solutions Chapter 8: Solutions, Acids, and Bases 8.1 Formation of Solutions 2. Stirring: moves dissolved particles away from surface of solid and allows more collisions 1. Conductivity NaCl HCl See pg. 231 Dispersion of Molecular Compounds 1. Surface Area high surface area=more collisions

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The Role of Enthalpy in Solution Formation. Energy is required to overcome the intermolecular interactions in a solute, which can be supplied only by the new interactions that occur in the solution, when each solute particle is surrounded by particles of the solvent in a process called solvation (or hydration when the solvent is water). Thus all of the solute–solute interactions and many of ...

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